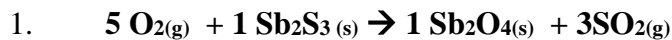


Name: _____ Per: _____

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Unit 9 Handout 1

Solve the following problems: Day 1: 1-3 all (1step problems) Day 2: 4 & 5 (1-3 step problems)

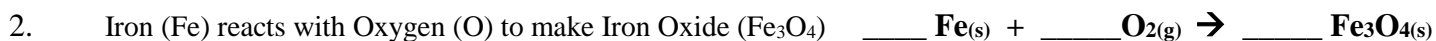


a) How many moles of Sb_2S_3 are needed to react completely with 2.18 O_2 moles?

b) How many SO_2 liters will be made from 14.9 liters of O_2 reacting?

c) 6.72 moles of O_2 would require how many moles of Sb_2S_3 in order to react?

d) To make 3.91 moles of Sb_2O_4 would require how many moles of Sb_2S_3 ?

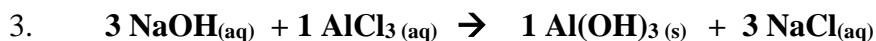


a) Balance the chemical reaction

b) How many Fe moles are needed to make 1.45 moles of Fe_3O_4 ?

c) 4.70 moles of Fe would need how many moles of O_2 to react?

d) 7.83 moles of Fe would make how many moles of Fe_3O_4 formula units?



a) How many moles of $\text{Al}(\text{OH})_3$ can be made from 2.71 moles of AlCl_3 ?

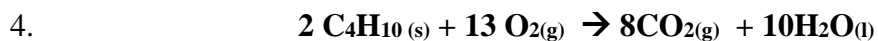
b) How many moles of AlCl_3 are needed for .0631 moles of NaOH to react?

c) To make .00697 moles of NaCl would require how much AlCl_3 ?

Name: _____ Per: _____

Date: _____

Day Two

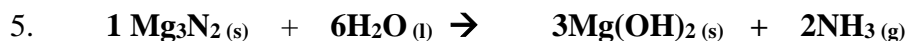


a) 3.58 liters of C_4H_{10} would need how many liters of O_2 to react with it?

b) 3.58 moles of C_4H_{10} would make how many grams of CO_2 ?

c) 14.72g of O_2 will make how many grams of CO_2 ?

d) To make 2816g of H_2O would require how many grams of C_4H_{10} ?



a) .470 moles of H_2O would make how many grams of $\text{Mg}(\text{OH})_2$?

b) If 49.2g of $\text{Mg}(\text{OH})_2$ were made, how many grams of NH_3 was also made?

c) 526g of Mg_3N_2 would need how many moles of H_2O to react with it?