

Chemistry: Unit 6 - Worksheet 6

More Practice with Names and Formulas

IONS			FORMULA	NAME
1. Na^+	and	Br^-	_____	_____
2. Cu^+	and	SO_4^{2-}	_____	_____
3. Pb^{2+}	and	Cl^-	_____	_____
4. K^+	and	S^{2-}	_____	_____
5. Sn^{2+}	and	F^-	_____	_____
6. _____	and	_____	BaI_2	_____
7. _____	and	_____	AlCl_3	_____
8. _____	and	_____	$\text{Sr}(\text{NO}_3)_2$	_____
9. _____	and	_____	KOH	_____
10. _____	and	_____	$(\text{NH}_4)_2\text{SO}_4$	_____
11. _____	and	_____	_____	silver oxide
12. _____	and	_____	_____	lithium bromide
13. _____	and	_____	_____	copper (II) nitrate
14. _____	and	_____	_____	manganese(II) chloride
15. _____	and	_____	_____	calcium carbonate
16. Mg^{2+}	and	NO_3^-	_____	_____
17. Cu^{2+}	and	OH^-	_____	_____
18. _____	and	_____	Na_3PO_3	_____
19. _____	and	_____	_____	iron (III) sulfide
20. _____	and	_____	_____	potassium carbonite

Part II

Write the names of the following compounds

1. $\text{Cu}(\text{NO}_3)_2$ _____
2. BaCl_2 _____
3. HgO _____
4. $\text{Ni}(\text{OH})_2$ _____
5. Na_3PO_4 _____
6. CaCO_3 _____
7. CS_2 _____
8. $\text{Sn}(\text{BrO}_3)_4$ _____
9. $(\text{NH}_4)_2\text{SeO}_4$ _____
10. $\text{Mg}(\text{NO}_3)_2$ _____
11. Li_2O _____
12. FeS _____
13. NI_3 _____
14. H_2SO_4 _____
15. KCN _____

Part III

Write the formulas for the following compounds (not all are ionic)

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|--------------------------------|---------------------------------|
| 1. copper (II) sulfate _____ | 2. sodium chromate _____ |
| 3. iron (III) chloride _____ | 4. silver sulfide _____ |
| 5. aluminum oxide _____ | 6. zinc nitrate _____ |
| 7. potassium phosphate _____ | 8. strontium fluoride _____ |
| 9. ammonium carbonate _____ | 10. magnesium hydroxide _____ |
| 11. carbon tetrachloride _____ | 12. phosphorus tribromide _____ |
| 13. sulfur hexafluoride _____ | 14. sulfur dioxide _____ |
| 15. chromium (III) oxide _____ | 16. nitric acid _____ |
| 17. hydrochloric acid _____ | 18. lead(II) iodide _____ |
| 19. ammonium nitrite _____ | 20. potassium perchlorate _____ |