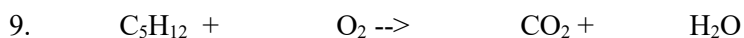
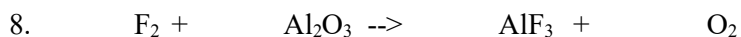
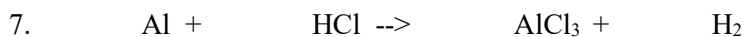
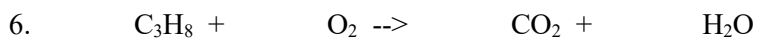
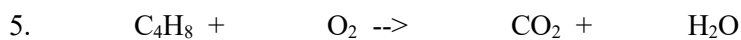
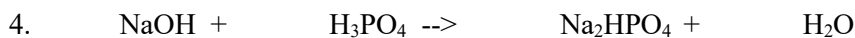
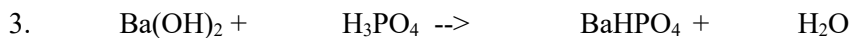


Name: _____ Per: _____

Unit 8 Extra Practice

Balance the following equations by filling in the correct coefficient.



Classify the above reactions

C— combustion	DR—double replacement	SR—single replacement	S—synthesis	D—decomposition
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Write equations for the following ions these dissolving in water You need to include the state of matter!!!

11) For Lithium Borate (3 & no#) **dissolving** in water

12) For Magnesium Permanganate (12 & no#) **dissolving** in water

Write equations for the following ions forming a precipitate when mixed in water.

You need to include the state of matter!!!

13) Write an equations for Aluminum Acetate (13 & no#) forming a precipitate when mixed in water.

14) Write an equations for Lithium phosphate (3& no#) forming a precipitate when mixed in water.

Write equations for the following reactions: You need to include the state of matter!!!

- 15) Solid Nickel metal reacts with solid phosphorus (P_4) to form solid nickel phosphide. Write a balanced equation.
- 16) Gaseous Carbon tetrachloride reacts with calcium metal producing solid carbon and calcium chloride. Write a balanced equation.
- 17) Aluminium metal reacts with liquid ammonium perchlorate to make solid aluminium oxide, chloride gas, water and nitrogen monoxide (laughing gas). Write a balanced equation.
- 18) Sulfur gas (S_8) reacts with solid copper and oxygen gas to make solid copper sulfate. Write a balance equation.
-

- 19) Octane (C_8H_{18}) and oxygen react in a combustion reaction. Predict the products and write a balanced equation.
-

Write net ionic equations for the following precipitates forming when 2 solutions are mixed. If no precipitate forms write no reaction (NR). You need to include the state of matter!!!

Rules: Alkali ions (Li^+ , Na^+ , K^+ , Rb^+ , and Cs^+), Nitrate (NO_3^{-1}) ions, Ammonium (NH_4^{+1}) and Acetate ($C_2H_3O_2^{-1}$) ions are **VERY SOLUBLE** with everything and stay in solution.

- 20) If solutions of cobalt (II) acetate (29&no#) and barium iodate (56&no#) are mixed a ppt forms. Write a net ionic equation for the ppt. formation.
- 21) If solutions of calcium iodate (20&no #) and lithium phosphate (3&no#) are mixed a ppt forms. Write a net ionic equation for the ppt. formation.
- 22) If solutions of Sodium carbonate (11&no #) and potassium permanganate (no#&no#) are mixed a ppt forms. Write a net ionic equation for the ppt. formation.